THE PREDICTION OF pK_a VALUES FOR PHENOLIC COMPOUNDS BY THE DFT THEORY

Nguyen Thi My, Nguyen Van Din, Mai Van Bay*

The University of Danang - University of Science and Education

*Corresponding author: mvbay@ued.udn.vn

(Received: January 25, 2022; Accepted: March 24, 2022)

Abstract - The acid dissociation constant is an important parameter that affects the physicochemical properties of molecules in solution. A set of 20 phenolic compounds were used to establish a model to predict pK_a values of phenolic compounds. Calculations in aqueous medium were performed with a polarizable continuum solvent model (PCM) and three hybrid DFT functionals (B3LYP, PBE0, ωB97XD and M062X) with the basis set 6-311++G(d,p). The directly calculated value of pK_a gives less accurate results with an average absolute error (MAE) of 1.74 pK_a units when using the ωB97XD functional and phenol as reference compound. In the case of using statistical correction, the accuracy of pK_a is greatly improved. In the case of using statistical correction, the accuracy of pK_a is greatly improved with the lowest MAE value of 0.14 pK_a units (M062X; $R^2 = 0.978$). The calculated results of pK_a in this study have the same accuracy as the experimental measurements.

Key words - pK_a; density functional theory methods (DFT); phenolic.

1. Introduction

The proton transfer reactions are one of the most basic and common types of reactions in chemistry [1]. The ability of a substance to transfer protons in a medium is characterized by its acid dissociation constant (K_a) and is usually reported as a pK_a value. The many chemical compounds act as Brønsted–Lowry acids or bases in aqueous media. Therefore, depending on the pK_a and pH values, these compounds can be ionized to different degrees and thus determine their forms of existence in the aquatic environment. The bioactive molecules from nature as well as drug molecules are usually weak acids or weak bases, so their degree of ionization in the medium affects lipophilicity, solubility, protein binding and mobility across the plasma membrane and thus pK_a affects the absorption, distribution, metabolism, excretion and toxicity properties of compounds [2]. Therefore, the pK_a value is a very important parameter for pharmaceutical compounds as well as other commonly used compounds. Most of the pK_a values are determined experimentally. However, the number of experimental pK_a values is too small compared to the extremely abundant number of chemical compounds. Therefore, there have been many publications using different theoretical approaches to predict pK_a value such as using the QSAR model (quantitative structure – activity relationship) combined with machine learning method [3, 4]; The semiempirical quantum mechanical methods and the density functional theory methods (DFT) [5-7]. The results of these publications show that it is possible to predict pK_a from the theory. However, each family of compounds needs to be approached separately, and the quantitative results of each model must be based on compounds with similar chemical structures.

The subjects in this study were phenolic compounds. They are compounds with one or more hydroxyl groups attached to the aromatic benzene nucleus and are found in most plant tissues [8]. Phenolics are known to have biological activities, such as antibacterial and antiviral properties, anti-inflammatory and antiproliferative activities and especially many phenolics have strong antioxidant activity and are the most abundant source of antioxidants in the human diet [9].

In aqueous media, phenolics are weak acids, which is due to the proton dissociation of the OH group attached to the benzene ring. Depending on the pK_a value and pH of the body's environment, phenolics can exist as neutral molecules or ionic form, or both. Therefore, the acid-base equilibrium of phenolics in the environment affects their bioactivity [10, 11] and pK_a values of phenolics are of the utmost importance in practical applications. In this study, we use DFT theory combined with a linear regression model to predict pK_a of phenolic compounds based on a sample of 20 phenolic compounds with known pK_a from the experiment (Figure 1).

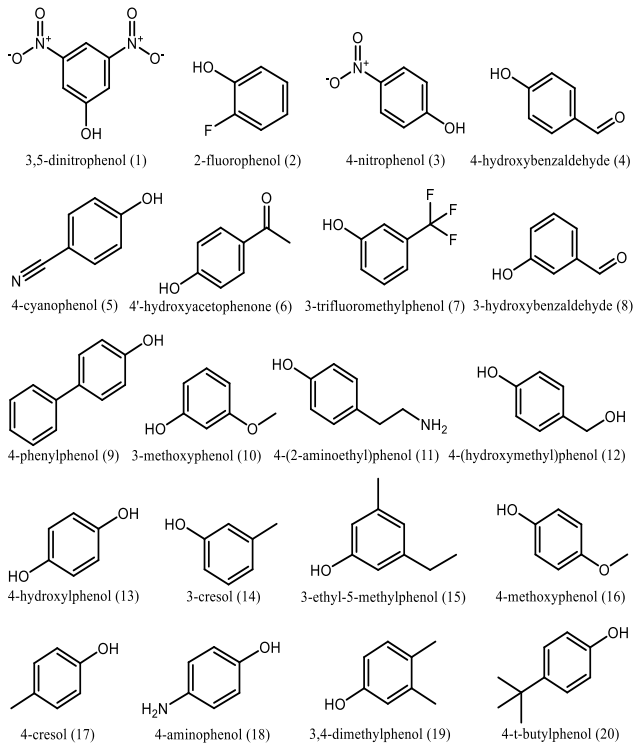

Figure 1. *Phenolic compounds used in this study*

2. Theoretical and computational methods

2.1. Calculation of pKa

The dissociation of phenolics (ArOH) in aqueous medium (aq) is expressed in terms of Equation (1).

$$\mathrm{ArOH_{(aq)} + H_2O_{(aq)} \rightleftharpoons ArO^-_{(aq)} + H_3O^+_{(aq)}} \quad (1)$$

Acid dissociation constant (K_a) and pK_a are defined according to equations (2) and (3) respectively[12].

$$K_a = [H_3O^+][ArO^-]/[ArOH] \quad (2)$$

$$pK_a = -\log K_a \quad (3)$$

The value of K_a is calculated based on Equation (4).

$$\Delta G_T^0(1) = -RT \ln K_a \quad (4)$$

Where, $\Delta G_T^0(1)$ is the change in Gibbs energy of acid dissociation, calculated by DFT theory. The disadvantage of this method is that it is difficult to accurately calculate the Gibbs energy of the H_3O^+ ion because the ion is very strongly hydrated in solution[13]. This can be overcome by using an acid reference ($ArOH_{ref}$). Consider the following equilibria:

$$\mathrm{ArOH_{(aq)} + ArO^-_{ref,(aq)} \rightleftharpoons ArO^-_{(aq)} + ArOH_{ref,(aq)}} \quad (5)$$

The pKa value of ArOH is determined according to Equation (6) (detailed description in section 1.1 of supplementary information).

$$pK_a = \frac{\Delta G_T^0(5)}{RT \ln 10} + pK_a^{ref} \quad (6)$$

Where, pK_a^{ref} is the experimental pK_a value the reference phenolics, $ArOH_{ref}$ and $\Delta G_T^0(5)$ is the change in Gibbs energy of reaction on Equation (5), determined by:

$$\Delta G_T^0(5) = G_T^0(ArOH_{ref}) + G_T^0(ArO^-) - G_T^0(ArOH) - G_T^0(ArO^-_{ref}) \quad (7)$$

The study used 3 reference phenolic compounds with significantly different pK_a values including phenol ($pK_a = 9.99$), 2-hydroxybenzaldehyde ($pK_a = 8.37$) and 2-cyanophenol ($pK_a = 6.86$). The predicted value of pK_a by direct calculation method (denoted by pK_a^{calc}) is calculated according to Equation (3).

The results are statistically processed as follows: from the pK_a^{calc} values, build a linear regression equation according to Equation (8).

$$pK_a^{exp} = a + b.pK_a^{calc} \quad (8)$$

Where, pK_a^{exp} is the experimental pK_a of phenolics used in the study sample. The final predicted pK_a after statistical correction (denoted by pK_a^{corr}) is calculated in terms of Equation (9).

$$pK_a^{corr} = \alpha + \beta.pK_a^{calc} \quad (9)$$

Where, α và β are equal to the values of a and b in Equation (8), respectively.

2.2. The DFT calculation

In this study, four density functional theory (DFT) methods including B3LYP [14], PBE0 [15], ωB97XD [16] and M06-2X [17] combinded with the 6-311++G(d,p) basic set were recommended for predict pK_a values. The choise of these methods was based on the previous studies because of its accuracy and lower computational cost [18-22]. In addition, the solvent effects were modeled by the polarizable continuum model (PCM) which is the most commonly used solvation model [23, 24].

The stable geometry of the compounds was checked by harmonic frequency calculation to confirm the structures which were minimal on the PES surface and to obtain the thermal and entropy contributions to the Gibbs energy. All calculations were performed at standard conditions in the solution of 1 M and 298.15 K and used Gaussian 16 Revision A.03 software[25].

3. Results and discussion

3.1. Direct calculation of pK_a

The pK_a error between calculation and experiment of 20 phenol compounds when using different DFT functionals and reference compounds is presented in Table S2 (supplementary information: SI) and visualized in Figure 2. The results show that the error pK_a is less dependent on the type of used functional but strongly depends on the reference compound. The 20 phenolic compounds in Figure 2 are numbered in order of increasing pK_a value (see Table S1 in SI). The errors tend to increase as the pK_a of the reference compound is further away from the pK_a of the calculated compound. Specifically, the reference compound 2-cyanophenol with a small pK_a (6.86) will cause large errors for the compound with a large pK_a (compound 7 to 20 have a pK_a between 8.98 and 10.39). In contrast, phenol with large pK_a (9.99) produces small errors for compounds with similar pK_a but large errors for compounds with small pK_a (compound 1 to 6 with pK_a of 6.69 to 8.05). The case of 2-hydroxybenzaldehyde has a pK_a (8.37) lower than phenol and greater than 2-cyanophenol, so it generally does not cause too large errors in the regions with large and small pK_a in the sample of the studied compounds. A general trend that does not depend on the functionals or the used reference compounds is that the error is mostly negative for substances with small pK_a and positive for compounds with large pK_a.

The results in Figure 3 show that the mean absolute error is the smallest when using phenol as the reference compound and the largest when using 2-cyanophenol. This is explained by the fact that the majority of phenolics have a pK_a closer to phenol than 2-cyanophenol. For the sample of compounds studied, the best calculation pK_a results were observed by the combination of ωB97XD functional and phenol as the reference compound with a MAE value of 1.74. However, the error between calculated and experimental pK_a varied widely from -5.86 (compound 1) to 2.02 (compound 18). The results here indicate that the selection of a reference compound is very important in the theoretical prediction of pK_a value. The closer the pK_a of the reference compound is to the one of interests, the more accurate the result. In general, the pK_a calculated by the direct method using the reference compound has a large error compared with the experimental method, and therefore caution should be exercised when using this method to predict the pK_a of phenolic compounds.

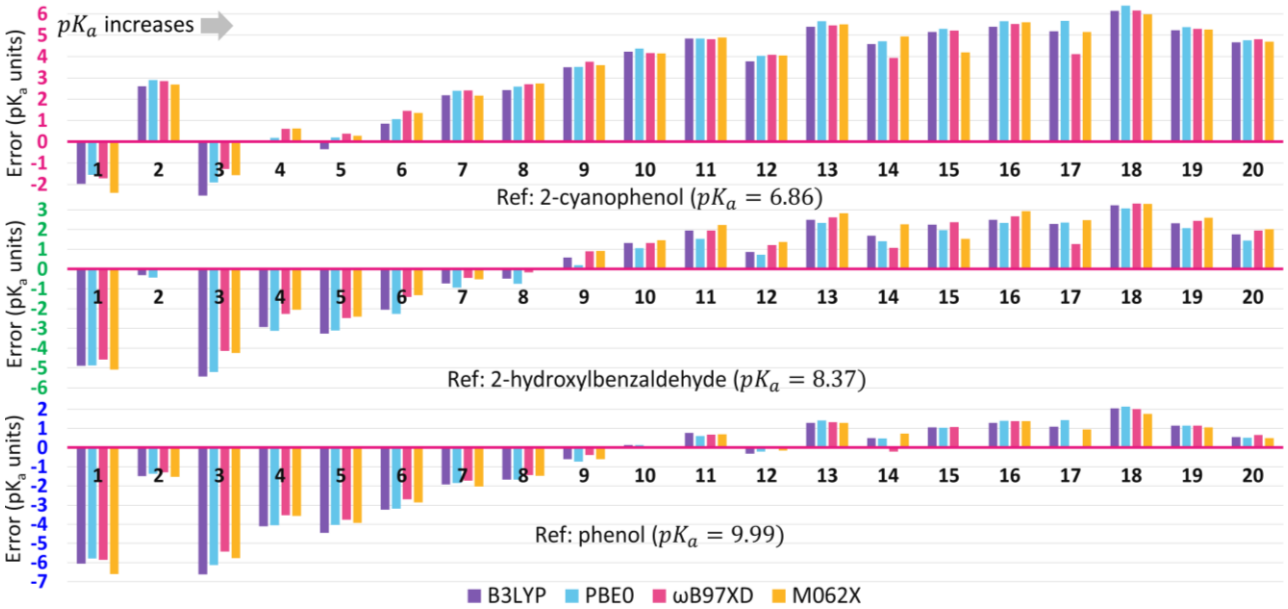

Figure 2. *The error differences between pK_a calculated and experimental values without statistical correction*

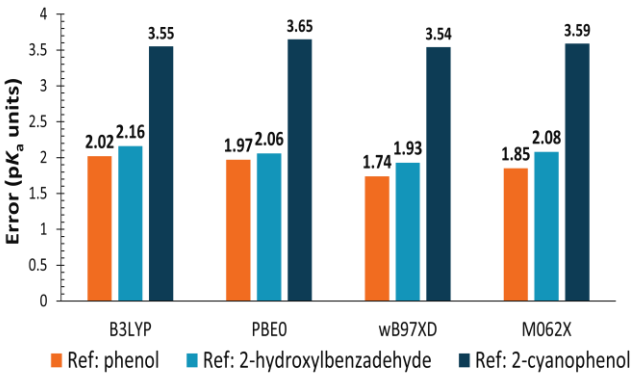

Figure 3. *The mean absolute error (MAE) differences between pKa calculated and experimental values without statistical correction*

3.2. The Calculated pK_a by statistical correction

The results in section 3.1 show that the pK_a when calculated by the direct method has a large uncertainty and is highly dependent on the reference compound. In this part of the study, we will examine the correlation between the calculated and experimental pK_a values. The results presented in Figure 4 show that there is a strong linear correlation between the calculated (pK_a^{calc}) and experimental pK_a (pK_a^{exp}) with the R^2 coefficient in the range from 0.972 to 0.978. Furthermore, the correlation coefficients and slopes in the regression equations do not depend on the type of reference compound selected. Therefore, the choice of the reference compound does not affect the degree of correlation between the calculated and experimental pK_a values. This means that the error in calculating pK_a after statistical correction will not depend on the type of reference compound and this will be analyzed below.

Based on the correlation between the calculated and experimental pK_a, the predicted pK_a value is statistically corrected (denoted by pK_a^{corr}) by replacing the quantity pK_a^{exp} in the regression equations with pK_a^{corr}. The equations for calculating pK_a^{corr} when using phenol reference compounds are presented in Table 1. Here, it should be noted that the choice of the reference compound is just a procedure when calculating, has absolutely no effect on the value of the pK_a^{corr} and is therefore the same when using the regression equations of the other reference compounds for the investigation.

Table 1. *Equations for calculation of pK_a^{corr} according to pK_a^{calc} of DFT functionals when phenol is reference*

Functionals	Equations	R^2
B3LYP	$pK_a^{corr} = 0.306 \times pK_a^{calc} + 6.715$	0.978
PBE0	$pK_a^{corr} = 0.314 \times pK_a^{calc} + 6.450$	0.978
ωB97XD	$pK_a^{corr} = 0.333 \times pK_a^{calc} + 6.617$	0.972
M062X	$pK_a^{corr} = 0.321 \times pK_a^{calc} + 6.099$	0.978

The error between pK_a^{corr} and pK_a^{exp} of 20 phenol compounds when using different DFT functionals is presented in Table S3 (SI) and visualized in Figure 5. The results show that the errors do not depend on the reference compound. Thus, unlike the case of direct pK_a calculation, the error here tends to be uniformly distributed throughout the sample where the pK_a varies over a wide range (from as low as 6.69 for compound 1 to a maximum of 10.39 for compound 20). The most important thing is that the error in calculating pK_a with statistical correction gives much better results than in the case of calculating pK_a directly. The maximum error of the studied phenolics was only 0.4 pK_a units for compound 17 when using the ωB97XD functional, while the largest error without statistical correction was 6.61 for compound 1 when using the M062X functional combined with phenol as the reference compound. As discussed, the error of the pK_a value after statistical correction is independent of the type of the used reference compound.

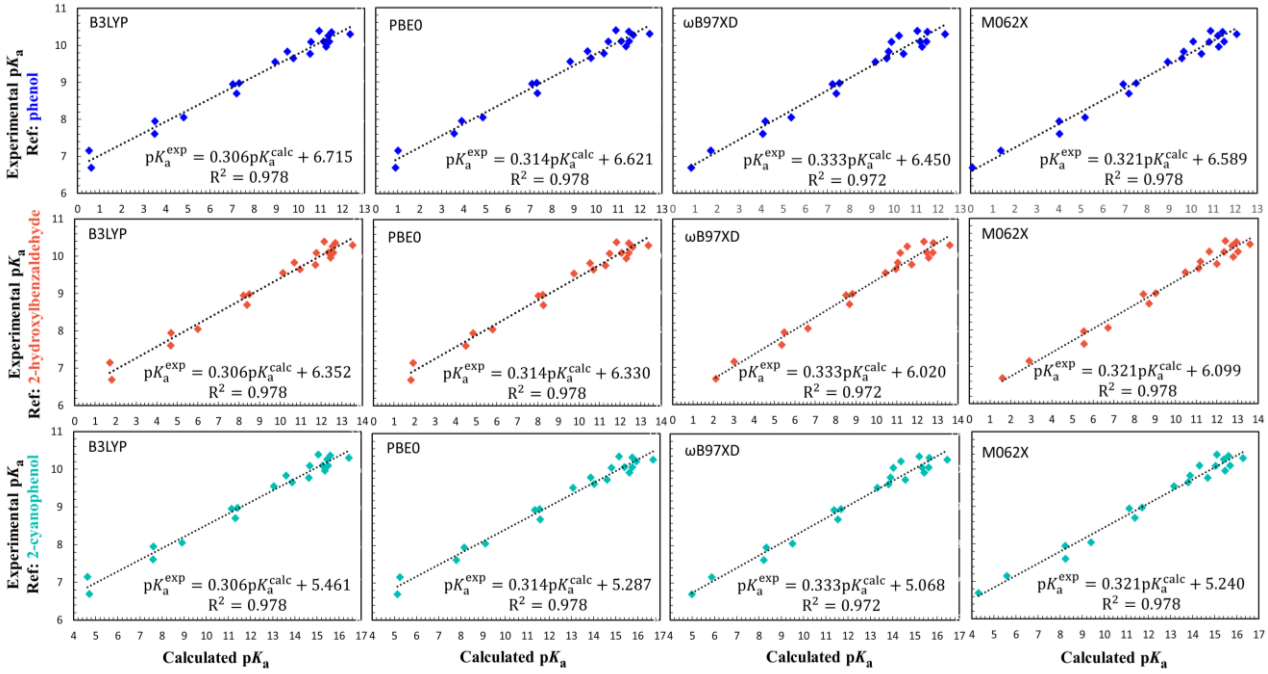

Figure 4. *Diagrams of linear correlation between the pK_a^{calc} and pK_a^{exp}*

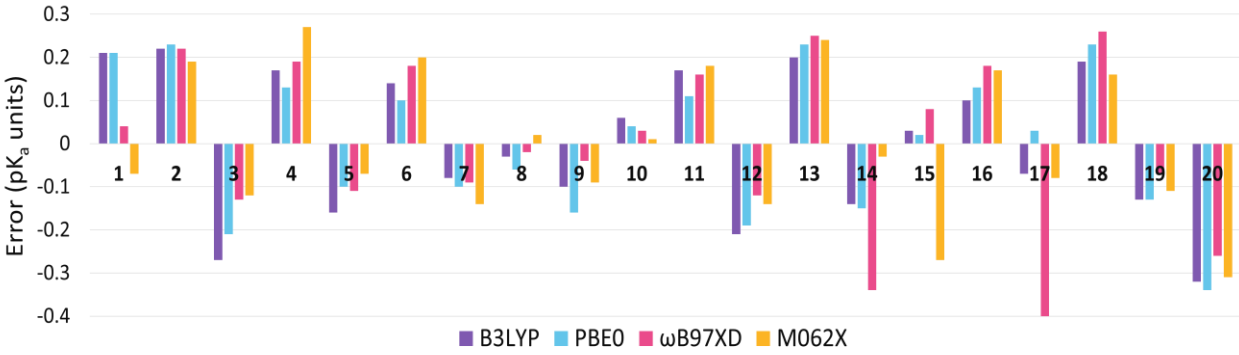

Figure 5. *The error differences between pK_a calculated and experimental values with statistical correction*

The results in Figure 6 show that there is a very small difference in the MAE values when using the B3LYP, PBE0, ωB97XD and M062X functionals. In which, the M062X functional gives the smallest MAE value of 0.14 pK_a units and the largest absolute error when using the M062X functional observed in compound 4 is 0.27 pK_a units. In summary, by using appropriate functionals combined with statistical correction, we are able to calculate the pK_a values of phenolic compounds with an accuracy asymptotic to the experimental measurements.

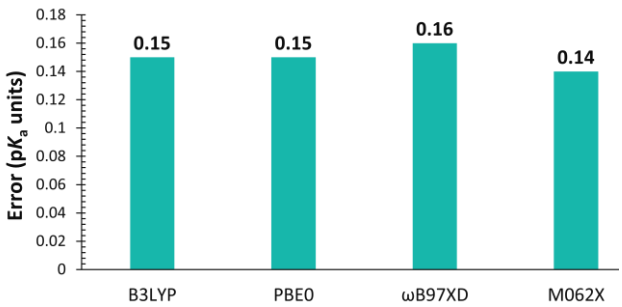

Figure 6. *The mean absolute error (MAE) differences between pK_a calculated and experimental with statistical correction*

4. Conclusions

This study presents a reliable procedure for calculating the pK_a of phenolic compounds in aqueous medium. The study sample consisted of 20 phenolic compounds and 3 reference phenolic compounds with experimentally known pK_a. Three hybrid DFT functionals (B3LYP, PBE0, ωB97XD and M062X) with the basis set 6-311++G(d,p) and the polarizable continuum solvent model (PCM) were used to calculate pK_a. The pK_a values calculated directly without statistical processing has a large error ($MAE \geq 1.74$ pK_a units) and the accuracy of the calculation depends strongly on the reference compound. The more accurate the result, the closer the reference compound has a pK_a to the compound of interest. In the sample of the compound studied, the ωB97XD functional combined with phenol as the reference compound gave the best results (MAE=1.74). (MAE = 1.74). The pK_a calculation results are greatly improved after statistical processing (MAE = 0.14 when using the M062X functional), with the error asymptotic with the experimental measurements. Furthermore, the pK_a calculation results are independent of the reference

compound. This is the recommended protocol for predicting pK_a values of phenolic compounds.

Acknowledgment: This research is funded by the University of Danang - Funds for Science and Technology Development under project number B2020-DN03-47.

REFERENCES

[1] M. Namazian and S. Halvani, "Calculations of pKa values of carboxylic acids in aqueous solution using density functional theory", *The Journal of Chemical Thermodynamics*, vol. 38, pp. 1495-1502, 2006.

[2] K. Mansouri, N. F. Cariello, A. Korotcov, V. Tkachenko, C. M. Grulke, C. S. Sprankle, *et al.*, "Open-source QSAR models for pKa prediction using multiple machine learning approaches", *Journal of cheminformatics*, vol. 11, pp. 1-20, 2019.

[3] J. Jover, R. Bosque, and J. Sales, "QSPR prediction of pKa for benzoic acids in different solvents", *QSAR & Combinatorial Science*, vol. 27, pp. 563-581, 2008.

[4] G. C. Shields and P. G. Seybold, *Computational approaches for the prediction of pKa values*: CRC Press, 2013.

[5] P. Hunt, L. Hosseini-Gerami, T. Chrien, J. Plante, D. J. Ponting, and M. Segall, "Predicting p K a Using a Combination of Semi-Empirical Quantum Mechanics and Radial Basis Function Methods", *Journal of chemical information and modeling*, vol. 60, pp. 2989-2997, 2020.

[6] R. Casasnovas, J. Ortega-Castro, J. Frau, J. Donoso, and F. Munoz, "Theoretical pKa calculations with continuum model solvents, alternative protocols to thermodynamic cycles", *International Journal of Quantum Chemistry*, vol. 114, pp. 1350-1363, 2014.

[7] J. H. Jensen, C. J. Swain, and L. Olsen, "Prediction of p K a Values for Druglike Molecules Using Semiempirical Quantum Chemical Methods", *The Journal of Physical Chemistry A*, vol. 121, pp. 699-707, 2017.

[8] C. Andrés-Lacueva, A. Medina-Remon, R. Llorach, M. Urpi-Sarda, N. Khan, G. Chiva-Blanch, *et al.*, "Phenolic compounds: chemistry and occurrence in fruits and vegetables", ed: Wiley Online Library, 2010, pp. 53-80.

[9] J. A. Vinson, Y. Hao, X. Su, and L. Zubik, "Phenol antioxidant quantity and quality in foods: vegetables", *Journal of agricultural and food chemistry*, vol. 46, pp. 3630-3634, 1998.

[10] J. Chen, J. Yang, L. Ma, J. Li, N. Shahzad, and C. K. Kim, "Structure-antioxidant activity relationship of methoxy, phenolic hydroxyl, and carboxylic acid groups of phenolic acids", *Scientific reports*, vol. 10, pp. 1-9, 2020.

[11] A. Aspée, C. Aliaga, L. Maretti, D. Zúñiga-Núñez, J. Godoy, E. Pino, *et al.*, "Reaction kinetics of phenolic antioxidants toward photoinduced pyranine free radicals in biological models", *The Journal of Physical Chemistry B*, vol. 121, pp. 6331-6340, 2017.

[12] R. A. Alberty, "Recommendations for nomenclature and tables in biochemical thermodynamics (IUPAC recommendations 1994)", *Pure and applied chemistry*, vol. 66, pp. 1641-1666, 1994.

[13] S. Zhao, Z. Jin, and J. Wu, "New theoretical method for rapid prediction of solvation free energy in water", *The Journal of Physical Chemistry B*, vol. 115, pp. 6971-6975, 2011.

[14] A. D. Becke, "A new mixing of Hartree–Fock and local density-functional theories", *The Journal of chemical physics*, vol. 98, pp. 1372-1377, 1993.

[15] J. P. Perdew, M. Ernzerhof, and K. Burke, "Rationale for mixing exact exchange with density functional approximations", *The Journal of chemical physics*, vol. 105, pp. 9982-9985, 1996.

[16] J.-D. Chai and M. Head-Gordon, "Long-range corrected hybrid density functionals with damped atom–atom dispersion corrections", *Physical Chemistry Chemical Physics*, vol. 10, pp. 6615-6620, 2008.

[17] Y. Zhao and D. G. Truhlar, "The M06 suite of density functionals for main group thermochemistry, thermochemical kinetics, noncovalent interactions, excited states, and transition elements: two new functionals and systematic testing of four M06-class functionals and 12 other functionals", *Theoretical chemistry accounts*, vol. 120, pp. 215-241, 2008.

[18] V. S. Bryantsev, "Predicting the stability of aprotic solvents in Li-air batteries: pKa calculations of aliphatic C–H acids in dimethyl sulfoxide", *Chemical Physics Letters*, vol. 558, pp. 42-47, 2013.

[19] F. Huang, J. Jiang, M. Wen, and Z.-X. Wang, "Assessing the performance of commonly used DFT functionals in studying the chemistry of frustrated Lewis pairs", *Journal of Theoretical and Computational Chemistry*, vol. 13, p. 1350074, 2014.

[20] R. Krishnan, J. S. Binkley, R. Seeger, and J. A. Pople, "Self-consistent molecular orbital methods. XX. A basis set for correlated wave functions", *The Journal of chemical physics*, vol. 72, pp. 650-654, 1980.

[21] S. Kheirjou, A. Abedin, A. Fattahi, and M. M. Hashemi, "Excellent response of the DFT study to the calculations of accurate relative pKa value of different benzo-substituted quinuclidines", *Computational and Theoretical Chemistry*, vol. 1027, pp. 191-196, 2014.

[22] F. M. Carvalho, Y. A. d. O. Só, A. S. K. Wernik, M. d. A. Silva, and R. Gargano, "Accurate acid dissociation constant (pKa) calculation for the sulfachloropyridazine and similar molecules", *Journal of Molecular Modeling*, vol. 27, pp. 1-9, 2021.

[23] S. Miertuš, E. Scrocco, and J. Tomasi, "Electrostatic interaction of a solute with a continuum. A direct utilizaion of AB initio molecular potentials for the prevision of solvent effects", *Chemical Physics*, vol. 55, pp. 117-129, 1981.

[24] A. Alibakhshi and B. Hartke, "Improved prediction of solvation free energies by machine-learning polarizable continuum solvation model", *Nature communications*, vol. 12, pp. 1-7, 2021.

[25] M. Frisch, G. Trucks, H. Schlegel, G. Scuseria, M. Robb, J. Cheeseman, *et al.*, "Gaussian 16 revision a. 03. 2016; gaussian inc", *Wallingford CT*, vol. 2, 2016.

[26] B. G. Tehan, E. J. Lloyd, M. G. Wong, W. R. Pitt, J. G. Montana, D. T. Manallack, *et al.*, "Estimation of pka using semiempirical molecular orbital methods. part 1: Application to phenols and carboxylic acids", *Quantitative Structure-Activity Relationships*, vol. 21, pp. 457-472, 2002.

[27] M. D. Liptak, K. C. Gross, P. G. Seybold, S. Feldgus, and G. C. Shields, "Absolute p K a determinations for substituted phenols", *Journal of the American Chemical Society*, vol. 124, pp. 6421-6427, 2002.

[28] C. Book, "Chemical Book",[Online]Available: http://www. chemicalbook. com/ProductChemicalPropertiesCB8852597_ EN. htm, 2017.

SUPPORTING INFORMATION (SI)

S1.1. The expression for calculating pKa based on reference compound

The dissociation equation for acid (phenolic: ArOH) in solution:

$$\mathrm{ArOH(aq) + H_2O(aq) \rightarrow ArO^-(aq) + H_3O^+(aq)} \quad (S1)$$

The dissociation equation for the reference phenolic ($ArOH_{ref}$) in solution:

$$\mathrm{ArOH_{ref}(aq) + H_2O(aq) \rightarrow ArO^-_{ref}(aq) + H_3O^+(aq)} \quad (S2)$$

Subtract Equation (S2) from Equation (S1) to get:

$$\mathrm{ArOH(aq) + ArO^-_{ref}(aq) \rightarrow ArO^-(aq) + ArOH_{ref}(aq)} \quad (S3)$$

Since Gibbs energy is a state function:

$$\Delta G^0_T(S3) = \Delta G^0_T(S1) - \Delta G^0_T(S2) \quad (S4)$$

Where, ΔG^0_T(Si) is the standard Gibbs energy of the reactions in terms of Eq (Si). The expression of standard Gibbs energy in terms of acid dissociation constant:

$$\Delta G^0_T(S1) = -RT \ln K_a \quad (S5)$$

$$\Delta G^0_T(S3) = -RT \ln K_a^{ref} \quad (S6)$$

Where, K_a and K_a^{ref} are the acid dissociation constant of ArOH and $\mathrm{ArOH_{ref}}$. Combining equations (S4), (S5) and (S6) get:

$$\Delta G^0_T(S3) = -RT \ln \frac{K_a}{K_a^{ref}} \quad (S7)$$

Equation (S7) can be rewritten:

$$-\log K_a = \frac{\Delta G^0_T(3)}{RT \ln 10} - \log K_a^{ref} \quad (S8)$$

To the definition $pK_a = -\log K_a$, finally get the expression pK_a in terms of pK_a^{ref}:

$$pK_a = \frac{\Delta G^0_T(5)}{RT \ln 10} + pK_a^{ref} \quad (S9)$$

S1.2. SUPPORT DATA

Table S1. *Sequence numbers of compounds and their experimental pK_a values*

No.	Compound	pK_a^{exp}	Ref
	2-cyanophenol	6.86	[24]
	2-hydroxylbenzaldehyde	8.37	[24]
	phenol	9.99	[25]
1	3,5-dinitrophenol	6.69	[24]
2	2-fluorophenol	6.86	[26]
3	4-nitrophenol	7.15	[24]
4	4-hydroxybenzaldehyde	7.61	[24]
5	4-cyanophenol	7.95	[24]
6	4'-Hydroxyacetophenone	8.05	[24]
7	3-trifluoromethylphenol	8.95	[25]
8	3-hydroxybenzaldehyde	8.98	[24]
9	4-phenylphenol	9.55	[24]
10	3-methoxyphenol	9.65	[25]
11	4-(2-aminoethyl)phenol	9.77	[24]
12	4-(hydroxylmethyl)phenol	9.83	[24]
13	4-hydroxylphenol	9.96	[24]
14	3-cresol	10.09	[25]
15	3-ethyl-5-methylphenol	10.1	[24]
16	4-methoxyphenol	10.1	[24]
17	4-cresol	10.26	[25]
18	4-aminophenol	10.3	[25]
19	3,4-dimethylphenol	10.36	[24]
20	4-t-butylphenol	10.39	[24]

Table S2. *The error between calculated and experimental pK values when using DFT functionals and 3 reference compounds*

Com.	$\Delta pK_a = K_a^{calc} - K_a^{exp}$												pK_a^{exp}
	phenol				2–hydroxylbenzaldehyde				2–cyanonphenol				
	B3LYP	PBE0	ωB97XD	M062X	B3LYP	PBE0	ωB97XD	M062X	B3LYP	PBE0	ωB97XD	M062X	
1	–6.07	–5.80	–5.86	–6.61	–4.89	–4.87	–4.57	–5.08	–1.98	-1.55	–1.71	–2.40	6.69
2	–1.49	–1.36	–1.30	–1.52	–0.30	–0.43	–0.02	0.01	2.61	2.89	2.84	2.69	6.86
3	–6.63	–6.14	–5.43	–5.77	–5.44	–5.21	–4.14	–4.25	–2.53	-1.90	–1.28	–1.57	7.15
4	–4.12	–4.05	–3.54	–3.57	–2.93	–3.13	–2.26	–2.05	–0.02	0.19	0.60	0.63	7.61
5	–4.45	–4.04	–3.77	–3.93	–3.26	–3.11	–2.48	–2.4	–0.35	0.21	0.38	0.28	7.95
6	–3.24	–3.19	–2.70	–2.86	–2.06	–2.26	–1.41	–1.33	0.85	1.06	1.45	1.35	8.05
7	–1.92	–1.85	–1.73	–2.03	–0.73	–0.93	–0.44	–0.51	2.18	2.39	2.41	2.17	8.95
8	–1.67	–1.67	–1.44	–1.47	–0.48	–0.74	–0.16	0.05	2.43	2.58	2.70	2.73	8.98
9	–0.60	–0.73	–0.38	–0.61	0.59	0.20	0.90	0.92	3.50	3.52	3.76	3.60	9.55
10	0.13	0.13	0.02	–0.07	1.31	1.06	1.31	1.46	4.22	4.37	4.17	4.14	9.65
11	0.76	0.60	0.67	0.7	1.94	1.52	1.95	2.22	4.85	4.84	4.81	4.90	9.77
12	–0.32	–0.21	–0.06	–0.16	0.86	0.72	1.22	1.37	3.77	4.03	4.08	4.05	9.83
13	1.30	1.41	1.32	1.30	2.49	2.34	2.61	2.83	5.40	5.66	5.46	5.51	9.96
14	0.49	0.48	–0.20	0.74	1.68	1.40	1.08	2.26	4.59	4.72	3.94	4.94	10.09
15	1.05	1.04	1.08	0.00	2.24	1.97	2.37	1.52	5.15	5.29	5.22	4.2	10.10
16	1.30	1.40	1.38	1.39	2.49	2.33	2.67	2.92	5.40	5.65	5.52	5.60	10.10
17	1.10	1.43	–0.03	0.95	2.28	2.35	1.26	2.47	5.19	5.67	4.12	5.15	10.26
18	2.05	2.14	2.02	1.77	3.23	3.07	3.31	3.30	6.14	6.39	6.16	5.98	10.30
19	1.14	1.14	1.15	1.06	2.32	2.07	2.44	2.59	5.23	5.38	5.29	5.27	10.36
20	0.56	0.51	0.66	0.49	1.75	1.44	1.95	2.02	4.66	4.76	4.81	4.70	10.39

Table S3. *The error between calculated pK values (after statistical correction) and experiment when using DFT functionals and 3 reference compounds*

Comp.	$\Delta pK_a = pK_a^{corr} - pK_a^{exp}$											
	phenol				2–hydroxylbenzaldehyde				2–cyanonphenol			
	B3LYP	PBE0	ωB97XD	M062X	B3LYP	PBE0	ωB97XD	M062X	B3LYP	PBE0	ωB97XD	M062X
1	0.21	0.21	0.04	–0.07	0.21	0.21	0.04	–0.07	0.21	0.21	0.04	–0.07
2	0.22	0.23	0.22	0.19	0.22	0.23	0.22	0.19	0.22	0.23	0.22	0.19
3	–0.27	–0.21	–0.13	–0.12	–0.27	–0.21	–0.13	–0.12	–0.27	–0.21	–0.13	–0.12
4	0.17	0.13	0.19	0.27	0.17	0.13	0.19	0.27	0.17	0.13	0.19	0.27
5	–0.16	–0.10	–0.11	–0.07	–0.16	–0.10	–0.11	–0.07	–0.16	–0.10	–0.11	–0.07
6	0.14	0.10	0.18	0.20	0.14	0.10	0.18	0.20	0.14	0.10	0.18	0.20
7	–0.08	–0.10	–0.09	–0.14	–0.08	–0.10	–0.09	–0.14	–0.08	–0.10	–0.09	–0.14
8	–0.03	–0.06	–0.02	0.02	–0.03	–0.06	–0.02	0.02	–0.03	–0.06	–0.02	0.02
9	–0.16	–0.04	–0.09	–0.10	–0.10	–0.16	–0.04	–0.09	–0.10	–0.16	–0.04	–0.09
10	0.06	0.04	0.03	0.01	0.06	0.04	0.03	0.01	0.06	0.04	0.03	0.01
11	0.17	0.11	0.16	0.18	0.17	0.11	0.16	0.18	0.17	0.11	0.16	0.18
12	–0.21	–0.19	–0.12	–0.14	–0.21	–0.19	–0.12	–0.14	–0.21	–0.19	–0.12	–0.14
13	0.20	0.23	0.25	0.24	0.20	0.23	0.25	0.24	0.20	0.23	0.25	0.24
14	–0.14	–0.15	–0.34	–0.03	–0.14	–0.15	–0.34	–0.03	–0.14	–0.15	–0.34	–0.03
15	0.03	0.02	0.08	–0.27	0.03	0.02	0.08	–0.27	0.03	0.02	0.08	–0.27
16	0.10	0.13	0.18	0.17	0.10	0.13	0.18	0.17	0.10	0.13	0.18	0.17
17	–0.07	0.03	–0.40	–0.08	–0.07	0.03	–0.40	–0.08	–0.07	0.03	–0.40	–0.08
18	0.19	0.23	0.26	0.16	0.19	0.23	0.26	0.16	0.19	0.23	0.26	0.16
19	–0.13	–0.13	–0.07	–0.11	–0.13	–0.13	–0.07	–0.11	–0.13	–0.13	–0.07	–0.11
20	–0.32	–0.34	–0.26	–0.31	–0.32	–0.34	–0.26	–0.31	–0.32	–0.34	–0.26	–0.31

Table S4. *The stable geometry (coordinates xyz) of the compounds is optimized at the theoretical level as M062X/6–311++G(d,p)/PCM(water)*

phenol				**phenol anion**			
Atom	x	y	z	Atom	x	y	z
C	0.267299	1.195992	–5E–06	C	0.289747	1.206222	0.000005
C	–1.12461	1.217336	0	C	–1.09832	1.19765	–6E–06
C	–1.85186	0.03077	0.000002	C	–1.81929	0	0
C	–1.17121	–1.18561	0	C	–1.09832	–1.19765	–6E–06
C	0.218104	–1.22054	0.000002	C	0.289747	–1.20622	0.000005
C	0.937017	–0.02626	0	C	1.061567	0	0.000046
H	0.75677	–2.1606	0.000006	H	0.828608	–2.14938	–1.1E–05
O	2.296897	–0.11329	–1.3E–05	O	2.344764	0	–2.5E–05
H	2.685609	0.767808	0.000105	H	–1.63184	–2.14412	–1.9E–05
H	–1.72487	–2.11729	0.000001	H	–1.63184	2.144117	–1.9E–05
H	–1.63875	2.171307	0.000001	H	0.828608	2.149382	–1.1E–05
H	0.832071	2.122422	–1.4E–05	H	–2.90241	0	–0.00001
H	–2.93444	0.052489	0.000004				
2–hydroxylbenzaldehyde				**2–hydroxylbenzaldehyde anion**			
Atom	x	y	z	Atom	x	y	z
C	1.270367	1.274811	0.000134	C	1.388297	1.196832	0.000009
C	2.298235	0.345707	–6.7E–05	C	2.307477	0.178887	0
C	2.036873	–1.02969	–0.00022	C	1.90815	–1.17773	–0.00006
C	0.724401	–1.4612	–0.00018	C	0.559912	–1.45121	–4.5E–05
C	–0.33457	–0.54105	0.000027	C	–0.42403	–0.43675	–0.00003
C	–0.05683	0.841369	0.000183	C	–0.04078	0.963579	–0.00013
O	–1.03041	1.766868	0.000402	O	–0.85438	1.92537	0.000146
H	–1.89162	1.307769	0.000447	H	0.220957	–2.4847	–2.6E–05
H	0.491167	–2.52106	–0.0003	H	3.366328	0.420646	0.000046
H	3.323847	0.696004	–0.0001	H	1.712102	2.232736	0.00009
H	1.470891	2.339104	0.000262	H	2.64296	–1.97246	–6.1E–05
H	2.851663	–1.74155	–0.00038	C	–1.79845	–0.89415	0.000086
C	–1.70885	–1.03201	0.000078	H	–1.88416	–2.00153	–4.4E–05
H	–1.82899	–2.12711	–0.00052	O	–2.82833	–0.23429	–1.7E–05
O	–2.69393	–0.30946	–0.00029				
2–cyanophenol				**2–cyanophenol anion**			
Atom	x	y	z	Atom	x	y	z
C	–0.46373	–0.40331	–4E–06	C	0.506422	–0.33181	–5.2E–05
C	0.416849	–1.49449	–2E–06	C	–0.32095	–1.46812	–1.1E–05
C	1.782194	–1.27945	0.000001	C	–1.69531	–1.34049	–1.4E–05
C	2.270757	0.030342	0.000004	C	–2.24691	–0.0431	0.000023
C	1.409742	1.115566	0.000003	C	–1.45437	1.08194	–0.00002
C	0.030814	0.910015	–1E–06	C	–0.01672	1.013106	–0.00023
H	1.781396	2.132459	0.000004	H	–1.8999	2.071357	0.000068
O	–0.76534	1.999128	0	O	0.721937	2.039342	0.000097
H	–1.70025	1.757923	–0.00002	H	–2.3329	–2.21477	0.000029
H	2.463892	–2.11946	0.000001	H	–3.32646	0.072565	0.000097
H	3.339749	0.205697	0.000006	H	0.140035	–2.45028	0.000039
H	0.009145	–2.49744	–5E–06	C	1.920584	–0.48977	0.000034
C	–1.87922	–0.60863	–1.3E–05	N	3.069616	–0.61775	0.000084
N	–3.02509	–0.7275	0.000014				
3,5–dinitrophenol (1)				**3,5–dinitrophenol anion**			
Atom	x	y	z	Atom	x	y	z
C	–1.21272	1.21492	0.000283	C	1.208852	1.259878	0.000112
C	–1.16602	–0.17025	0.000259	C	1.165987	–0.1196	0.000009
C	0.007446	–0.90011	–6.7E–05	C	–5E–06	–0.87432	–3.5E–05
C	1.174314	–0.15433	–0.00037	C	–1.16599	–0.11959	0.000016
C	1.202365	1.227701	–0.00035	C	–1.20885	1.25989	0.000125
C	–0.00992	1.917923	–2.5E–05	C	0.000005	2.034772	0.000203
H	2.138855	1.768238	–0.00059	O	0.000017	3.298134	0.000179
O	0.054213	3.263217	–2.7E–05	H	2.160054	1.7747	0.00012
H	–0.82652	3.655666	0.000217	H	–1.1E–05	–1.95141	–0.00013
H	–2.16658	1.726307	0.000542	N	2.449211	–0.85737	–0.00006
N	–2.4463	–0.91013	0.000593	N	–2.44921	–0.85736	–7.3E–05
N	2.463179	–0.87958	–0.00072	O	3.480892	–0.2175	–0.00031
O	–3.47048	–0.26297	0.000007	O	2.410126	–2.07165	0.000097
O	–2.39628	–2.12041	–0.00066	O	–3.4809	–0.21749	0.000355
O	3.48014	–0.2216	0.000319	O	–2.41014	–2.07163	–0.00055
O	2.426075	–2.09049	0.000663	C	1.208852	1.259878	0.000112
C	1.208852	1.259878	0.000112				
2–fluorophenol (2)				**2–fluorophenol anion**			
Atom	x	y	z	Atom	x	y	z
C	–0.48189	–0.69802	0.00001	C	–0.55975	–0.63912	–9E–06
C	0.71598	–1.38095	–0.00004	C	0.595554	–1.38583	0.000021
C	1.900965	–0.64883	–7.6E–05	C	1.83825	–0.74165	0.000054
C	1.853409	0.742794	–5.9E–05	C	1.855973	0.652317	0.000045
C	0.633707	1.412442	–8E–06	C	0.676483	1.390478	0.000014
C	–0.55531	0.691425	0.000027	C	–0.6174	0.787806	–2E–06
H	0.583323	2.494551	0.000006	H	0.712078	2.475713	0.000002
O	–1.74368	1.345261	0.000075	O	–1.71853	1.437143	–6.7E–05
H	–2.46824	0.708175	0.000101	H	2.806552	1.176421	0.000061
H	2.772075	1.316088	–8.7E–05	H	0.516818	–2.46794	0.000019
H	0.707985	–2.46381	–5.1E–05	H	2.754844	–1.31731	0.000077
H	2.851534	–1.16614	–0.00012	F	–1.75297	–1.30534	–0.00004
F	–1.65538	–1.3738	0.000047				
4–nitrophenol (3)				**4–nitrophenol anion**			
Atom	x	y	z	Atom	x	y	z
C	1.380415	–1.20223	0.000036	C	–1.41593	1.224659	0.000049
C	–0.00331	–1.21104	0.000125	C	–0.04827	1.221332	0.000105
C	–0.68374	0.000714	0.000158	C	0.655813	0	–5.1E–05

C	–0.01468	1.221494	0.000088	C	–0.04827	–1.22133	–0.00012
C	1.366757	1.226605	–4E–06	C	–1.41593	–1.22466	–0.00018
C	2.066928	0.01572	–0.00003	C	–2.18911	0	–0.0004
H	1.930702	–2.13564	0.000017	H	–1.96212	2.161265	0.000279
H	1.920936	2.156206	–5.8E–05	H	–1.96212	–2.16127	–0.00012
O	3.412645	0.086423	–0.00012	O	–3.44395	0	0.000263
H	3.802924	–0.79557	–0.00013	H	0.50891	–2.14952	–0.00012
H	–0.57436	2.146549	0.00011	H	0.50891	2.149523	0.000283
H	–0.55198	–2.14266	0.000176	N	2.067659	0	0
N	–2.14319	–0.0077	0.000257	O	2.672161	1.076207	0.000201
O	–2.71296	–1.08287	–0.00025	O	2.672161	–1.07621	–5.6E–05
O	–2.7247	1.061131	–0.00015				

4–hydroxybenzaldehyde (4) **4–hydroxybenzaldehyde anion**

Atom	x	y	z	Atom	x	y	z
C	1.209812	1.186794	–1E–06	C	1.253863	1.197874	–0.0001
C	–0.16978	1.330689	0.000003	C	–0.1135	1.32923	–8.5E–05
C	–1.00178	0.210065	0.000007	C	–0.96597	0.204337	–0.00023
C	–0.43481	–1.07113	0.000006	C	–0.37075	–1.07836	–5.9E–05
C	0.936677	–1.22769	0.000003	C	0.991055	–1.23103	–6.8E–05
C	1.761339	–0.09535	–1E–06	C	1.889877	–0.0969	–0.00065
H	1.395203	–2.20879	0.000003	H	1.438434	–2.21972	0.000239
O	3.096183	–0.3105	–6E–06	O	3.144434	–0.23641	0.000627
H	3.574613	0.526384	–4E–06	H	1.896087	2.072304	0.000191
H	1.857911	2.05605	–5E–06	H	–0.55946	2.321321	0.000048
H	–0.60582	2.324363	0.000003	H	–1.01962	–1.94893	0.000081
H	–1.08669	–1.9369	0.000009	C	–2.39463	0.38459	–0.00031
C	–2.46064	0.393312	0.000012	H	–2.72416	1.442284	0.000297
H	–2.80029	1.445079	–0.00002	O	–3.2408	–0.5043	0.000395
O	–3.26866	–0.5103	–1.4E–05				

4–cyanophenol (5) **4–cyanophenol anion**

Atom	x	y	z	Atom	x	y	z
C	–1.01819	–1.20085	0.000024	C	1.060103	1.21662	–2E–06
C	0.366615	–1.20842	0.000005	C	–0.3148	1.21246	–9E–06
C	1.071509	–0.0013	–1.4E–05	C	–1.03397	–1.8E–05	0.0001
C	0.379093	1.215837	–1.5E–05	C	–0.31478	–1.21248	0.00002
C	–1.00331	1.223901	0.000004	C	1.060124	–1.21662	0.000026
C	–1.7043	0.015432	0.000023	C	1.832777	0.000009	0.000333
H	–1.55596	2.154782	0.000003	H	1.601924	–2.1564	–0.00014
O	–3.0536	0.086238	0.00004	O	3.097804	0.00002	–0.00039
H	–3.4409	–0.79663	0.000059	H	–0.86022	–2.15049	–0.00006
H	0.92764	2.149335	–3.1E–05	H	–0.86026	2.150455	–0.00011
H	0.902784	–2.14905	0.000006	H	1.601886	2.156409	–0.00019
H	–1.5683	–2.13476	0.000039	C	–2.45425	–3.5E–05	0.000075
C	2.504053	–0.01002	–3.4E–05	N	–3.61099	0.000034	0.000048
N	3.65582	–0.01729	–0.00005				

4'–hydroxyacetophenone (6) **4'–hydroxyacetophenone anion**

Atom	x	y	z	Atom	x	y	z
C	–1.49694	–1.18142	–0.00003	C	1.541957	1.196946	0.000052
C	–0.11381	–1.23932	–0.00006	C	0.169436	1.248036	–0.0001
C	0.656148	–0.0718	–0.00006	C	–0.62155	0.077809	–4.9E–05
C	0.002477	1.165687	–1.7E–05	C	0.056784	–1.15999	0.000042
C	–1.38125	1.23822	0.000018	C	1.431214	–1.23281	0.000204
C	–2.13312	0.06309	0.000011	C	2.259722	–0.05462	0.000464
H	–2.08786	–2.09112	–3.9E–05	H	2.127994	2.110604	–7.9E–05
H	0.389551	–2.19846	–9.1E–05	H	–0.33785	2.207381	–0.00028
H	–1.89374	2.192112	0.000052	H	1.930128	–2.19646	0.000185
O	–3.48124	0.188369	0.000044	O	3.52411	–0.11209	0.000056
H	–3.90017	–0.67965	0.000047	H	–0.51174	–2.0846	–1.9E–05
H	0.570782	2.087825	–0.00001	C	–2.07804	0.181486	–0.00019
C	2.142371	–0.18596	–9.7E–05	C	–2.89832	–1.09506	–0.00011
C	2.963256	1.080238	0.000002	H	–3.95526	–0.83674	–0.00023
H	4.019275	0.820419	0.000031	H	–2.67251	–1.70012	–0.88073
H	2.733813	1.683292	–0.88129	H	–2.67266	–1.69991	0.880696
H	2.73375	1.683186	0.881351	O	–2.65853	1.265718	–0.00023
O	2.681218	–1.27712	0.000124				

3–trifluoromethylphenol (7) **3–trifluoromethylphenol anion**

Atom	x	y	z	Atom	x	y	z
C	–2.42285	0.816182	0	C	–2.47733	0.748347	0
C	–1.47795	1.833437	0	C	–1.55603	1.780708	0.000001
C	–0.11665	1.541426	0	C	–0.17899	1.532906	–1E–06
C	0.275724	0.208486	–1E–06	C	0.22657	0.199997	–2E–06
C	–0.65317	–0.82632	–1E–06	C	–0.67637	–0.85529	–3E–06
C	–2.00896	–0.51705	0	C	–2.08881	–0.63317	–6E–06
H	–3.4816	1.051905	0.000001	H	–3.54057	0.968618	0.000003
H	–1.80698	2.86506	0	H	–1.90853	2.807241	0.000004
H	–0.34321	–1.86492	–1E–06	H	–0.31907	–1.88065	–3E–06
O	–2.88238	–1.55589	0	O	–2.93573	–1.58813	0.000002
H	–3.78991	–1.23259	0	H	0.540344	2.339918	0
H	0.619369	2.333783	–1E–06	C	1.683813	–0.14082	0
C	1.729144	–0.16175	0	F	2.042402	–0.87613	–1.07541
F	2.059506	–0.90114	–1.07569	F	2.042404	–0.87611	1.075421
F	2.059509	–0.90113	1.075699	F	2.483701	0.93821	–1.1E–05
F	2.537605	0.905307	–9E–06				

4–hydroxybenzaldehyde (8) **4–hydroxybenzaldehyde anion**

Atom	x	y	z	Atom	x	y	z
C	1.209812	1.186794	–1E–06	C	1.253863	1.197874	–0.0001
C	–0.16978	1.330689	0.000003	C	–0.1135	1.32923	–8.5E–05
C	–1.00178	0.210065	0.000007	C	–0.96597	0.204337	–0.00023
C	–0.43481	–1.07113	0.000006	C	–0.37075	–1.07836	–5.9E–05
C	0.936677	–1.22769	0.000003	C	0.991055	–1.23103	–6.8E–05
C	1.761339	–0.09535	–1E–06	C	1.889877	–0.0969	–0.00065
H	1.395203	–2.20879	0.000003	H	1.438434	–2.21972	0.000239
O	3.096183	–0.3105	–6E–06	O	3.144434	–0.23641	0.000627
H	3.574613	0.526384	–4E–06	H	1.896087	2.072304	0.000191
H	1.857911	2.05605	–5E–06	H	–0.55946	2.321321	0.000048
H	–0.60582	2.324363	0.000003	H	–1.01962	–1.94893	0.000081
H	–1.08669	–1.9369	0.000009	C	–2.39463	0.38459	–0.00031
C	–2.46064	0.393312	0.000012	H	–2.72416	1.442284	0.000297
H	–2.80029	1.445079	–0.00002	O	–3.2408	–0.5043	0.000395
O	–3.26866	–0.5103	–1.4E–05				

4–Phenylphenol (9) **4–Phenylphenol anion**

Atom	x	y	z	Atom	x	y	z
C	2.410493	–1.13045	0.386516	C	–2.46498	1.157543	0.335881
C	1.020686	–1.12724	0.385272	C	–1.08097	1.14881	0.330453
C	0.296951	0.004565	–0.00037	C	–0.3425	–7.2E–05	0.000139
C	1.015462	1.142193	–0.38732	C	–1.081	–1.14887	–0.33042
C	2.402754	1.152479	–0.39181	C	–2.46501	–1.1576	–0.33572
C	3.104425	0.01264	–0.00387	C	–3.24013	–0.00019	0.000658
H	2.95451	–2.01569	0.698206	H	–3.00043	2.061174	0.61151
H	0.494144	–2.01746	0.71052	H	–0.55289	2.054263	0.616337
H	2.954076	2.032074	–0.70173	H	–3.00048	–2.06103	–0.61194
O	4.464836	0.074121	–0.02594	O	–4.51908	0.000168	–0.00052
H	4.840495	–0.76872	0.25026	H	–0.55295	–2.05422	–0.61667
H	0.482152	2.029027	–0.71098	C	1.13551	–4.2E–05	0.000052
C	–1.1871	0.00021	0.000211	C	1.864396	–1.15197	0.336489
C	–1.90679	1.134176	0.3944	C	1.864268	1.151945	–0.33646
C	–1.90015	–1.13811	–0.39362	C	3.255207	–1.15412	0.330913
C	–3.29776	1.130763	0.39325	H	1.333704	–2.05184	0.626784
H	–1.37341	2.019066	0.723644	C	3.255081	1.154205	–0.33103
C	–3.29108	–1.14339	–0.39124	H	1.333466	2.05177	–0.62668
H	–1.36153	–2.01943	–0.7239	C	3.961468	0.000069	–0.0001
C	–3.99593	–0.00853	0.001458	H	3.789738	–2.05852	0.599222
H	–3.83676	2.016828	0.707821	H	3.78951	2.058647	–0.5994
H	–3.82481	–2.03274	–0.70554	H	5.044921	0.000112	–0.00017
H	–5.0794	–0.01186	0.001886	C	–2.46498	1.157543	0.335881
C	2.410493	–1.13045	0.386516	C	–1.08097	1.14881	0.330453
C	1.020686	–1.12724	0.385272	C	–0.3425	–7.2E–05	0.000139
C	0.296951	0.004565	–0.00037	C	–1.081	–1.14887	–0.33042
C	1.015462	1.142193	–0.38732				

3–Methoxyphenol (10) **3–Methoxyphenol anion**

Atom	x	y	z	Atom	x	y	z
C	1.731493	1.028845	–5E–06	C	1.790703	0.988005	0.000144
C	0.563515	1.777573	0.000003	C	0.638164	1.750552	0.000227
C	–0.69368	1.174835	0.000019	C	–0.64219	1.178058	0.000177
C	–0.76705	–0.22021	0.000022	C	–0.71719	–0.21878	–1.1E–05
C	0.397252	–0.99055	0.000007	C	0.429547	–1.00747	–9.1E–05
C	1.635519	–0.36528	–4E–06	C	1.740438	–0.44739	0.000054
H	0.329751	–2.07137	0.000002	H	0.319336	–2.08694	–0.00026
O	2.735662	–1.1686	–2E–06	O	2.798764	–1.17106	–0.00021
H	3.535496	–0.63169	–0.00013	H	2.765898	1.464336	0.000158
H	2.702611	1.51057	–1.1E–05	H	0.720194	2.833494	0.000328
H	0.6281	2.859211	0.000002	H	–1.52237	1.804138	0.000243
H	–1.58326	1.787983	0.000031	O	–1.90451	–0.90745	–0.00011
O	–1.92992	–0.92076	0.000043	C	–3.10165	–0.15361	–0.0001
C	–3.14488	–0.18687	–4.6E–05	H	–3.91658	–0.87434	–0.00023
H	–3.9425	–0.92524	–8.2E–05	H	–3.1738	0.475508	0.891577
H	–3.22466	0.437687	0.893185	H	–3.17368	0.475698	–0.89166
H	–3.22454	0.437655	–0.89331				

4–(2–aminoethyl)phenol (11) **4–(2–aminoethyl)phenol anion**

Atom	x	y	z	Atom	x	y	z
C	1.704712	1.189525	0.021236	C	1.754461	1.198654	0.011355
C	0.342369	1.196794	0.306608	C	0.395918	1.19361	0.300137
C	–0.37797	0.010411	0.447066	C	–0.32996	0.005562	0.443917
C	0.311199	–1.19505	0.289681	C	0.380841	–1.18913	0.281961
C	1.670837	–1.22163	0.004699	C	1.739204	-1.20704	-0.00743
C	2.369883	-0.02436	-0.13077	C	2.502381	-0.00776	-0.16278
H	2.24945	2.122518	-0.0793	H	2.285918	2.141092	-0.08641
H	-0.16693	2.148029	0.424894	H	-0.12196	2.143087	0.421946
H	-0.22519	-2.13281	0.39611	H	-0.14873	-2.13382	0.389989
H	2.201231	-2.15936	-0.1113	H	2.258665	–2.15452	–0.11976
O	3.704126	–0.10084	–0.40858	O	3.759984	–0.01355	–0.43076
H	4.078246	0.784175	–0.47519	C	–1.81692	0.01531	0.697887
C	–1.86369	0.03036	0.702333	H	–2.10428	–0.87299	1.272566
H	–2.15108	–0.84724	1.29059	H	–2.09776	0.887719	1.296367
H	–2.13754	0.915485	1.282501	C	–2.63039	0.039002	–0.59717
C	–2.66156	0.034386	–0.60303	H	–2.32813	–0.81576	–1.21773
H	–2.35645	–0.83154	–1.20629	H	–2.38324	0.946312	–1.15577
H	–2.40564	0.931429	–1.17368	N	–4.06878	0.056841	–0.30563
N	–4.09914	0.063573	–0.32042	H	–4.33246	–0.81389	0.147032
H	–4.37651	–0.80683	0.124243	H	–4.59967	0.094076	–1.16998
H	–4.62334	0.115202	–1.1879				

4–(hydroxylmethyl)phenol (12) **4–(hydroxylmethyl)phenol anion**

Atom	x	y	z	Atom	x	y	z
C	1.195917	1.193772	0.004572	C	1.21561	1.20646	–0.01896
C	–0.18303	1.228527	–0.18017	C	–0.15665	1.194852	–0.21341
C	–0.92272	0.055158	–0.30785	C	–0.87983	–2E–06	–0.31493
C	–0.2485	–1.16686	–0.24442	C	–0.15664	–1.19485	–0.2134
C	1.125308	–1.22027	–0.05877	C	1.215613	–1.20646	–0.01895
C	1.849149	–0.03458	0.066207	C	1.980409	0.000001	0.090656
H	1.651033	–2.16647	–0.01351	H	1.748839	–2.15016	0.049338

O	3.195814	–0.13957	0.243038	O	3.249284	0.000002	0.268227
H	3.590746	0.736411	0.308533	H	1.748835	2.150163	0.049333
H	1.762077	2.114455	0.098142	H	-0.68975	2.139943	-0.29248
H	-0.6872	2.188102	-0.22759	H	-0.68975	-2.13995	-0.29247
H	-0.80866	-2.09088	-0.34492	C	-2.36738	-3E-06	-0.46961
C	-2.41679	0.094239	-0.46104	H	-2.69149	0.888132	-1.02278
H	-2.72517	1.059707	-0.87358	H	–2.69149	–0.88815	–1.02277
H	–2.74489	–0.69602	–1.14364	O	–2.98686	0.000005	0.829147
O	–3.00491	–0.10138	0.8274	H	–3.94139	–1E–06	0.704419
H	–3.96123	–0.11773	0.721941				

4–hydroxylphenol (13) **4–hydroxylphenol anion**

Atom	x	y	z	Atom	x	y	z
C	0.7173	1.190085	0	C	0.760605	1.197961	0.000104
C	–0.67286	1.216374	0.000002	C	–0.62985	1.211812	–3E–06
C	–1.39414	0.027018	0	C	–1.35049	0.021658	0.000048
C	–0.7173	–1.19009	–4E–06	C	–0.65635	–1.18361	–2.4E–05
C	0.672858	–1.21637	0	C	0.736545	–1.20115	0.00011
C	1.394139	–0.02701	0	C	1.522061	–0.01054	0.000674
H	1.205192	–2.15985	0	H	1.256685	–2.15427	–0.00016
O	2.759533	–0.11337	0.000001	O	2.812537	–0.0229	–0.00045
H	3.144958	0.768888	0.000005	H	1.302993	2.138732	–0.00016
H	1.274122	2.121055	0	H	–1.17038	2.153412	–0.00018
H	–1.20519	2.15985	0.000005	H	–1.20744	–2.12091	–0.00023
H	–1.27413	–2.12105	–0.00001	O	–2.73513	0.093438	–0.00017
O	–2.75953	0.113368	–0.00001	H	–3.09625	–0.79806	0.000178
H	–3.14495	–0.7689	0.000088				

3–cresol (14) **3–cresol anion**

Atom	x	y	z	Atom	x	y	z
C	1.319554	1.017981	–0.00019	C	–1.38923	0.963063	0.000257
C	0.161501	1.786894	–5.3E–05	C	–0.25604	1.76073	0.000065
C	–1.09384	1.184503	0.000214	C	1.028517	1.205735	–0.00028
C	–1.20636	–0.20754	0.000312	C	1.162313	–0.18748	–0.00042
C	–0.04401	–0.97905	0.0003	C	0.025642	–0.99171	–0.00039
C	1.207993	–0.37216	–2.6E–05	C	–1.30362	–0.46745	0.000113
H	–0.10005	–2.06266	0.000531	H	0.137331	–2.07384	–0.00074
O	2.299419	–1.19048	–5.4E–05	O	–2.34416	–1.2201	0.000047
H	3.105879	–0.66405	–0.00007	H	–2.37808	1.412532	0.000403
H	2.297961	1.486558	–0.00034	H	–0.36851	2.841481	0.000074
H	0.24229	2.867719	–7.5E–05	C	2.531858	–0.82703	0.00034
C	–2.55391	–0.88263	–0.00025	H	2.672291	–1.45279	0.885561
H	–2.67156	–1.51237	–0.88512	H	2.665652	–1.46962	–0.87371
H	–2.6659	–1.52582	0.875563	H	3.320444	–0.07292	–0.0096
H	–3.36076	–0.14966	0.007734	H	1.907458	1.840764	–0.00053
H	–1.98878	1.796163	0.000399				

3–ethyl–5–methylphenol (15) **3–ethyl–5–methylphenol anion**

Atom	x	y	z	Atom	x	y	z
C	–0.36591	–1.20239	–0.23195	C	0.330765	1.263224	–0.2273
C	–0.95795	0.052436	–0.36412	C	0.934216	0.01812	–0.35991
C	–0.16438	1.19557	–0.24486	C	0.167064	–1.14818	–0.24162
C	1.205773	1.101344	0.002437	C	–1.20732	–1.05356	0.00911
C	1.786406	–0.16049	0.130935	C	–1.80208	0.195617	0.145995
C	1.002372	–1.3032	0.013608	C	–1.06729	1.417909	0.033834
H	2.849519	–0.26715	0.318743	H	–2.86905	0.264636	0.344715
O	1.628373	–2.50915	0.144736	O	–1.62243	2.569882	0.158313
H	0.990382	–3.22315	0.040879	H	0.930588	2.165976	–0.32085
H	–0.96808	–2.10234	–0.32197	H	0.638103	–2.12209	–0.34379
H	–0.62336	2.174159	–0.34877	C	–2.04638	–2.30782	0.094222
C	2.060846	2.336643	0.126542	H	–2.95134	–2.13962	0.680893
H	2.547839	2.374432	1.103676	H	–2.35524	–2.63705	–0.90247
H	2.848187	2.341259	–0.63076	H	–1.48684	–3.12785	0.549293
H	1.464623	3.241294	0.005885	C	2.428353	–0.08513	–0.57239
C	–2.44771	0.16974	–0.57469	H	2.646495	–0.92599	–1.23762
H	–2.66238	1.074255	–1.1492	H	2.791309	0.819718	–1.06768
H	–2.80172	–0.67977	–1.16427	C	3.183437	–0.27908	0.747935
C	–3.20289	0.218446	0.758754	H	2.852677	–1.19216	1.24854
H	–2.87702	1.07493	1.353323	H	4.260748	–0.35203	0.581431
H	–4.27884	0.304105	0.594882	H	2.997445	0.560784	1.42175
H	–3.0155	–0.68747	1.33975				

4–methoxyphenol (16) **4–methoxyphenol anion**

Atom	x	y	z	Atom	x	y	z
C	1.325919	1.164601	–3E–06	C	1.378173	1.161723	–6.5E–05
C	–0.04518	1.371953	0.000004	C	0.00684	1.365822	–2E–06
C	–0.92669	0.289315	0.000021	C	–0.88098	0.28789	0.000129
C	–0.41629	–1.00828	0.000029	C	–0.36029	–1.00408	0.000163
C	0.963079	–1.21291	0.000014	C	1.021693	–1.20908	0.000085
C	1.836079	–0.134	–2E–06	C	1.964402	–0.14304	0.000001
H	1.365444	–2.21881	0.000014	H	1.404714	–2.22524	0.000088
O	3.178564	–0.39835	–6E–06	O	3.241397	–0.32982	–0.00013
H	3.676846	0.425723	–0.00016	H	2.046492	2.017605	–0.00018
H	2.001155	2.013632	–1.2E–05	H	–0.40036	2.372386	–4.6E–05
H	–0.45044	2.376679	0	H	–1.01559	–1.86712	0.000246
H	–1.07154	–1.86864	0.000043	O	–2.22841	0.59815	0.000262
O	–2.25589	0.59614	0.000037	C	–3.13014	–0.48678	–0.00029
C	–3.17502	–0.48249	–4.9E–05	H	–3.00362	–1.11039	0.890772
H	–3.05394	–1.10192	0.892883	H	–4.13063	–0.0587	–0.00047
H	–4.16642	–0.03639	–8.5E–05	H	–3.00307	–1.10991	–0.89161
H	–3.05384	–1.10185	–0.89301				

4–cresol (17) **4–cresol anion**

Atom	x	y	z	Atom	x	y	z
C	–0.7515	–1.19151	–0.00067	C	0.789177	1.201294	–0.00242
C	0.642275	–1.19698	–0.00363	C	–0.60038	1.189914	–0.01175
C	1.376097	–0.01233	–0.0038	C	–1.33649	0.000039	–0.013
C	0.665067	1.192351	–0.00358	C	–0.6004	–1.18988	–0.01175
C	–0.72293	1.217238	–0.00069	C	0.78913	–1.20129	–0.00242
C	–1.43586	0.019215	0.001308	C	1.563811	0.000005	0.004547
H	–1.26517	2.155473	–0.00141	H	1.322679	–2.14768	–0.00607
O	–2.79901	0.095574	0.003134	O	2.850624	–4.2E–05	0.012141
H	–3.17907	–0.78915	0.003105	H	1.322735	2.147681	–0.00608
H	–1.30166	–2.12678	–0.00151	H	–1.13368	2.138231	–0.02009
H	1.162851	–2.14889	–0.00652	H	–1.13372	–2.13819	–0.0201
H	1.208871	2.131777	–0.00649	C	–2.84585	–2.2E–05	0.016924
C	2.883306	–0.01782	0.004685	H	–3.23541	–0.00285	1.04041
H	3.271831	0.382383	0.944709	H	–3.25074	0.884355	–0.48052
H	3.272117	–1.02965	–0.11586	H	–3.2509	–0.88157	–0.48542
H	3.283579	0.59921	–0.80287				

4–aminophenol (18) **4–aminophenol anion (18)**

Atom	x	y	z	Atom	x	y	z
C	–0.72407	–1.18861	–0.00068	C	0.766452	1.197245	–0.00045
C	0.667509	–1.19993	–0.00699	C	–0.62554	1.193053	–0.00935
C	1.39529	–0.00771	–0.00885	C	–1.35066	–2.8E–05	–0.01091
C	0.687824	1.199529	–0.00726	C	–0.6255	–1.1931	–0.00947
C	–0.70092	1.21413	–0.00031	C	0.766481	–1.19724	–0.00046
C	–1.41395	0.019032	0.002167	C	1.541759	0.000022	0.002911
H	–1.2422	2.152935	0.004249	H	1.296783	–2.14522	0.005379
O	–2.78303	0.092681	0.009362	O	2.834445	0.000028	0.012558
H	–3.15775	–0.79406	0.012091	H	–1.16237	–2.1387	–0.0142
H	1.232524	2.137602	–0.01117	H	–1.16242	2.138655	–0.01389
H	1.193228	–2.14869	–0.01051	H	1.296713	2.145254	0.005433
H	–1.27066	–2.12596	0.00355	N	–2.77605	–0.00016	–0.08927
N	2.798259	–0.01908	–0.08037	H	–3.17495	–0.81865	0.355389
H	3.232542	0.805047	0.313669	H	–3.17492	0.819828	0.352631
H	3.218599	–0.85339	0.307336				

3,4–dimethylphenol (19) **3,4–dimethylphenol anion**

Atom	x	y	z	Atom	x	y	z
C	–1.27768	–1.25256	0	C	–1.32323	–1.2502	–7.6E–05
C	0.075912	–1.58292	–2E–06	C	0.028663	–1.575	–1.5E–05
C	1.071085	–0.60922	–4E–06	C	1.035398	–0.6067	–9.2E–05
C	0.686134	0.744507	–3E–06	C	0.632984	0.742792	–8E–06
C	–0.66547	1.077533	–1E–06	C	–0.71813	1.074459	–8.5E–05
C	–1.64558	0.08709	0	C	–1.7692	0.10486	–0.00055
H	–2.03475	–2.02938	0.000001	H	–2.07212	–2.03713	0.000159
H	0.359574	–2.63025	–4E–06	H	0.315792	–2.62446	0.000128
H	–0.97481	2.117221	–1E–06	H	–1.00608	2.123252	0.000143
O	–2.94974	0.494861	0.000003	O	–3.0143	0.433417	0.000365
H	–3.53322	–0.27133	–5E–06	C	2.49506	–0.98859	–2.6E–05
C	2.528174	–0.99235	0.000003	H	3.021656	–0.60027	0.878243
H	3.044257	–0.59583	0.878976	H	2.607597	–2.07446	–4.2E–05
H	2.64351	–2.07652	–0.00005	H	3.021745	–0.60023	–0.87822
H	3.044292	–0.59574	–0.87891	C	1.677856	1.832733	0.000178
C	1.731184	1.82889	0	H	2.326897	1.759564	0.878133
H	2.377323	1.749966	0.878606	H	2.327066	1.759728	–0.87767
H	2.377323	1.749973	–0.87861	H	1.215322	2.820521	0.000228
H	1.27188	2.817198	0.000004				

4–t–butylyphenol (20) **4–t–butylyphenol anion**

Atom	x	y	z	Atom	x	y	z
C	–1.76299	–1.18119	–0.00002	C	–1.81073	–1.1921	–4E–06
C	–0.37408	–1.16185	–4.9E–05	C	–0.42445	–1.1637	–1.1E–05
C	0.351464	0.035563	–6.4E–05	C	0.308764	0.034176	–1.3E–05
C	–0.38454	1.222848	–6E–06	C	–0.44726	1.211572	0
C	–1.77736	1.223098	0.000004	C	–1.83979	1.205006	0
C	–2.47056	0.019346	–7E–06	C	–2.60202	–0.0001	–3E–06
H	–2.29691	–2.12586	–4E–06	H	–2.33188	–2.14562	0.000001
H	0.149129	–2.11194	–1.3E–05	H	0.105984	–2.113	–3E–06
H	0.121633	2.179809	–2E–06	H	0.052797	2.174585	0.000004
H	–2.33265	2.153716	0.000021	H	–2.38087	2.147304	0.000006
O	–3.83501	0.06825	0.000009	O	–3.88836	–0.01928	0.000004
H	–4.19644	–0.82431	0.000007	C	1.839132	0.004273	0.000001
C	1.881719	–0.00105	0.000002	C	2.441553	1.413079	0.000271
C	2.490218	1.404708	0.000927	H	2.140164	1.977597	0.886683
H	2.195072	1.970187	0.888621	H	2.140079	1.977972	–0.88587
H	2.194644	1.971526	–0.88578	H	3.532663	1.344282	0.000206
H	3.58001	1.326367	0.000619	C	2.348496	–0.73254	1.25123
C	2.37575	–0.74266	1.253841	H	1.977888	–1.75968	1.28628
H	2.007512	–1.77048	1.28375	H	3.442495	–0.76767	1.255952
H	3.46886	–0.77564	1.263053	H	2.016799	–0.22072	2.158833
H	2.039238	–0.23282	2.160432	C	2.348539	–0.73207	–1.25149
C	2.375977	–0.74108	–1.25469	H	1.977997	–1.75922	–1.2869
H	2.008048	–1.76897	–1.28582	H	2.016812	–0.21996	–2.15891
H	2.039359	–0.23027	–2.16069	H	3.44254	–0.76713	–1.25622
H	3.469096	-0.77374	–1.2639				